



Name \_\_\_\_\_

Chemistry Quiz 1: LEQ 1

Date \_\_\_\_\_

**Use this word bank for questions 1-6.**

**Atomic mass number      Element      Nucleus      Electron**

1. \_\_\_\_\_: The average mass of an atoms in an element (protons + neutrons)
2. \_\_\_\_\_: The center of an atom, it contains the protons and neutrons.
3. \_\_\_\_\_: A pure substance. All the atoms in it have the same atomic number or number of protons.
4. \_\_\_\_\_: Smallest subparticle with a negative charge.

**Use this word bank for questions 7-11.**

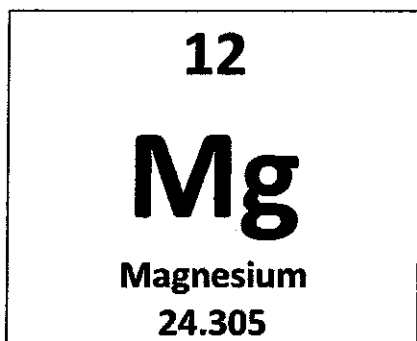
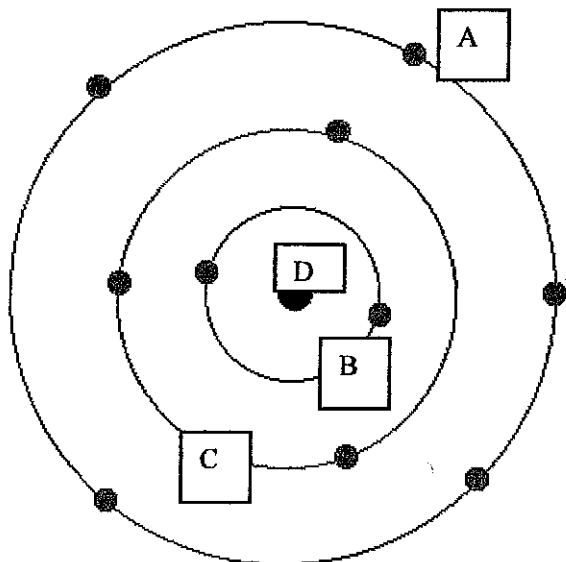
**Atomic Number      Proton      Neutron      Atoms**

5. \_\_\_\_\_: Large subparticle with no charge, it is neutral.
6. \_\_\_\_\_: The building blocks of matter. Made up of protons, electrons, and neutrons.
7. \_\_\_\_\_: Large subparticle with a positive charge.
8. \_\_\_\_\_: The number of protons or electrons found in an element.
9. \_\_\_\_\_
10. What makes an atom's charge electrically neutral?
  - a. subatomic particles have no charge
  - b. atoms contain only neutrons, which have no charge
  - c. the number of negative electrons equals the number of positive protons

**10. Use the picture of the atom to answer question #10 below**

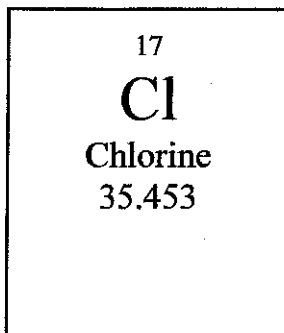
Which subatomic particles are found in the area of the atom marked D?

- a. neutrons and electrons
- b. protons and neutrons
- c. electrons and protons
- d. molecules and protons



11. The number "12" in the square above is:
- The number of electrons an electrically-neutral atom of magnesium contains
  - The number of protons an atom of magnesium contains
  - Magnesium's atomic number
  - All of the above are correct

Use the element box to answer questions 12-16.



- What is the **atomic number** of Chlorine? \_\_\_\_\_
- What is the **atomic mass** of Chlorine? \_\_\_\_\_
- How many **protons** are in Chlorine? \_\_\_\_\_
- How many **electrons** are in Chlorine? \_\_\_\_\_
- How many **neutrons** are in Chlorine? \_\_\_\_\_

Directions: Use the data table below to answer question number 17

17. Which element represents an atom of Phosphorous (P)?

- a. Element 1
- b. Element 2
- c. Element 3
- d. Element 4

Element	Protons	Neutrons	Electrons
Element 1	15Protons	16 Neutrons	16 Electrons
Element 2	31 Protons	16 Neutrons	31Electrons
Element 3	15 Protons	31 Neutrons	1 Electrons
Element 4	15 Protons	16 Neutrons	15Electrons

*LEQ1, Day 1: Atoms are the smallest form of matter*

Periodic table



Elements



Atoms

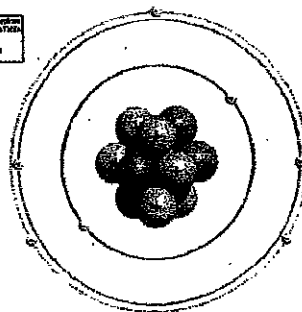
7

N

Nitrogen

14

Atomic Number  
= 7



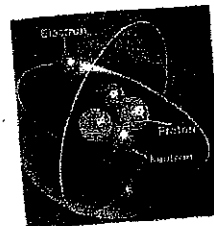
1. What are atoms?

• Atoms are the \_\_\_\_\_ form of \_\_\_\_\_ that make up everything in the world!

2. What are atoms made of?

• Atoms are made of 3 basic subatomic particles:  
\_\_\_\_\_, \_\_\_\_\_, & \_\_\_\_\_.

3. Nucleus: \_\_\_\_\_, where the protons and neutrons hang out.



4. The parts of an atom:

Subatomic Particle	Charge	Location	Symbol
Proton			
Neutron			
Electron			

Label the parts of an atom:

